

NAME \_\_\_\_\_  
 CHEMICAL REACTIONS TEST REVIEW

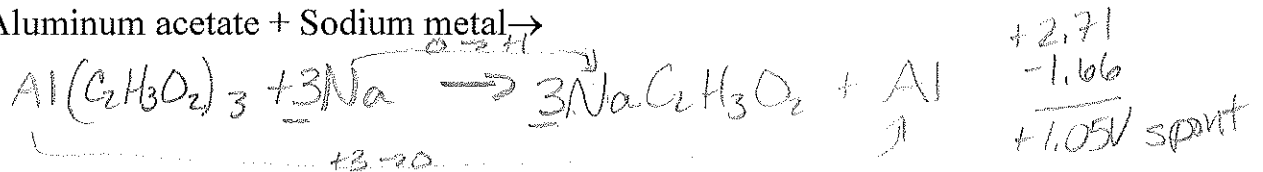
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**INSTRUCTIONS:**

Translate chemical name to symbols, indicate type of reaction, predict products (if necessary), and balance the equation.

Redox

1. Aluminum acetate + Sodium metal →



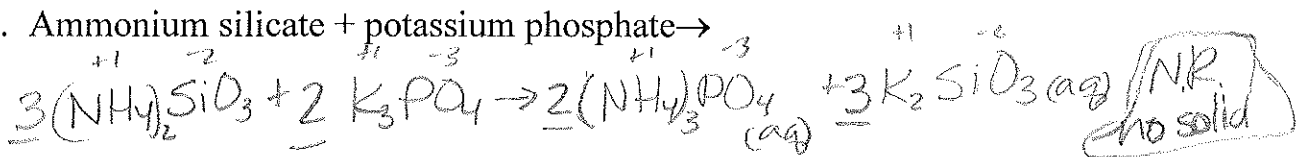
Sol

2. Cobalt(II) sulfide + Nickel(II) chloride →



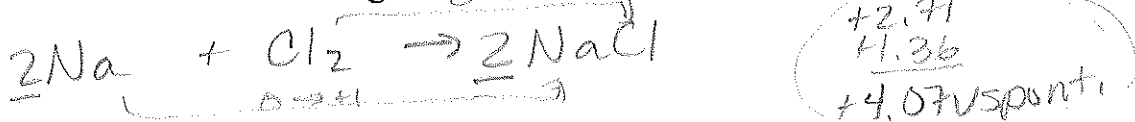
Sol

3. Ammonium silicate + potassium phosphate →



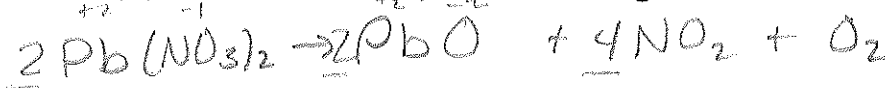
Redox  
Composition

4. sodium metal + Chlorine gas →



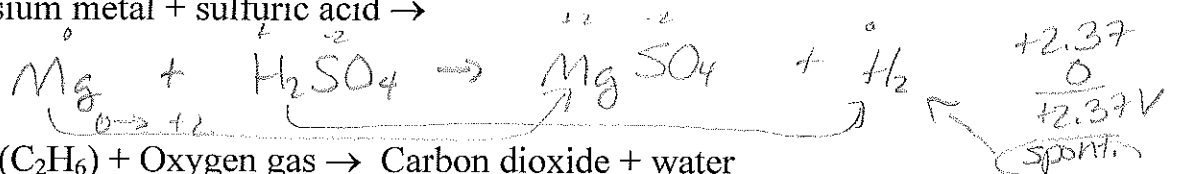
Redox  
Decomp

5. Lead (II) nitrate → lead(II) oxide + nitrogen dioxide + oxygen gas



Redox

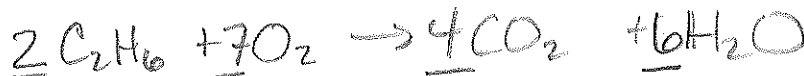
6. Magnesium metal + sulfuric acid →



Redox  
Combustion

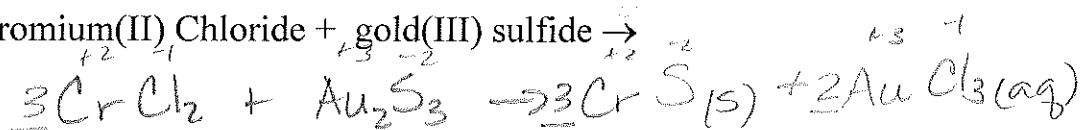
7. Ethane (C<sub>2</sub>H<sub>6</sub>) + Oxygen gas → Carbon dioxide + water

(combustion reaction, always result in carbon dioxide and water)



Sol

8. Chromium(II) Chloride + gold(III) sulfide →

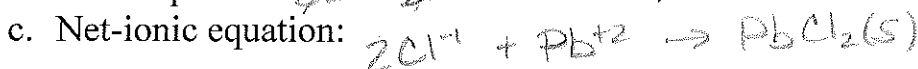
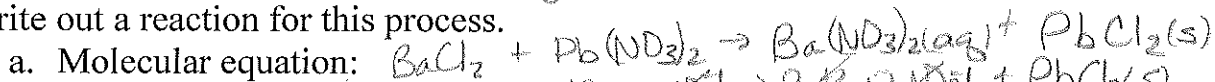


9. A container has the aqueous solutions poured into it and a reaction takes place causing a solid to be formed. ( $\text{BaCl}_{2(aq)} + \text{Pb}(\text{NO}_3)_2$ )

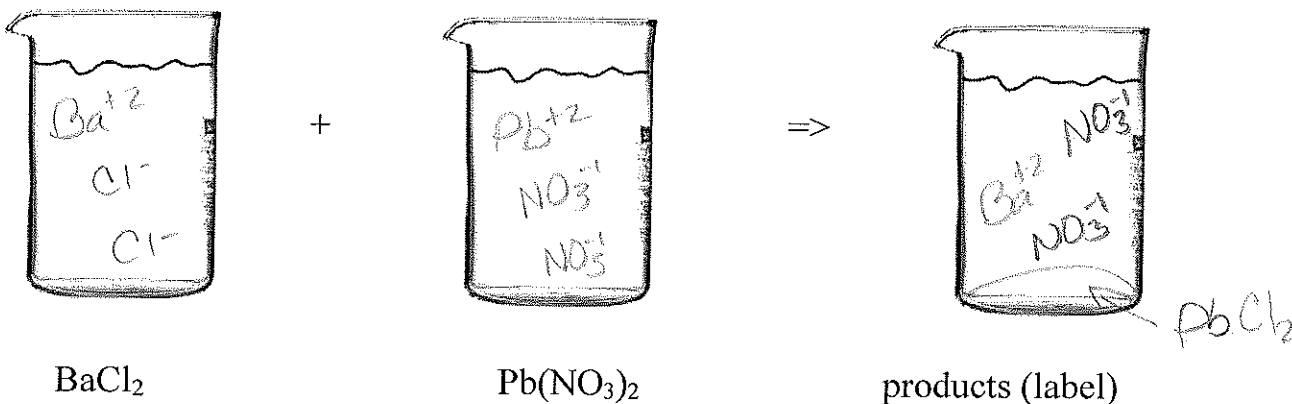
a. What type of chemical reaction is this?

*solubility*

b. Write out a reaction for this process.



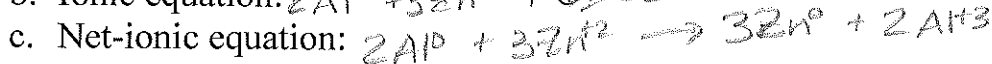
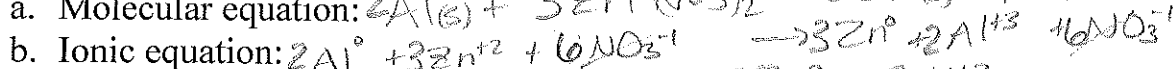
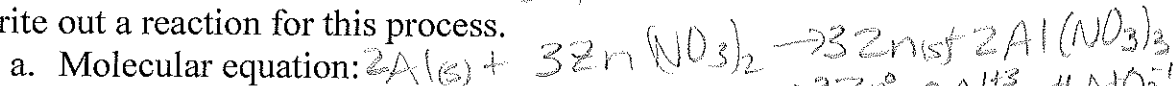
c. Draw a picture of each substance before and after.



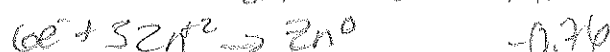
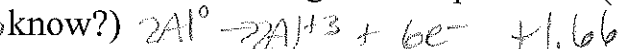
10. A solid chunk of aluminum is placed in zinc nitrate.

a. What type of reaction is this? *Redox*

b. Write out a reaction for this process.



c. Determine the voltage of this process? (is it spontaneous, how do you



$+0.9V$

*Check voltage*

d. Model this process before and after.

$\text{NO}_3^- = (\text{A})$   
(spectator)

